









Curriculum Overview for Chemistry Year 11.3 and 11.4

<p>Half Term 1: Chemical analysis</p> <p>Substantive Knowledge:</p> <ul style="list-style-type: none"> <input type="checkbox"/> Describe how we test for pure substances. <input type="checkbox"/> Describe the use of formulation. <input type="checkbox"/> Identify examples of formulations. <input type="checkbox"/> Describe the chromatography required practical. <input type="checkbox"/> Explain how paper chromatography separates mixtures <input type="checkbox"/> Describe the test for hydrogen. <input type="checkbox"/> Describe the test for oxygen <input type="checkbox"/> Describe the test for chlorine. <input type="checkbox"/> Describe the test for carbon dioxide <p>Disciplinary Knowledge:</p> <ul style="list-style-type: none"> <input type="checkbox"/> Calculate Rf values, <input type="checkbox"/> Interpret chromatograms <input type="checkbox"/> 		<p>Skim and Scan of source information Decoding terms Etymology of key terms Breaking down exam questions</p>
		<p>Chromatogram, Chromatography, spectroscopy, Formulation Impure substance, Litmus paper, Mobile phase, Precipitation, Pure substances, Rf value, Stationary phase</p>
		<p>Baseline Recall questions to start every lesson Exam questions in homework End of unit assessment</p>
		<p>Revision Card preparation for every lesson Exam questions - application</p>
<p>Half Term 2: The rate and extent of chemical change</p> <p>Substantive Knowledge:</p> <ul style="list-style-type: none"> ▪ Describe how changing temperature, concentration, pressure, surface area affects rate of reaction. ▪ Explain using collision theory the effects of changing conditions temperature, concentration, pressure, surface area on rate of reaction. ▪ Describe how catalysts impact rate of reaction. ▪ Link how catalysts impact rate of reaction to reaction profiles. ▪ Identify the symbol for reversible reactions. ▪ Link exothermic and endothermic reactions to reversible reactions. <p>Disciplinary Knowledge:</p> <ul style="list-style-type: none"> ▪ Calculate mean rate of reaction. ▪ Draw and interpret graphs showing the quantity of product formed or reactant used. ▪ 		<p>Skim and Scan of source information Decoding terms Etymology of key terms Breaking down exam questions</p>
		<p>Activation energy, Catalyst, Collision theory, Rate of reaction, Reversible reaction, concentrate</p>
		<p>Baseline Recall questions to start every lesson Exam questions in homework End of unit assessment</p>
		<p>Revision Card preparation for every lesson Exam questions - application</p>